

Isotopes 1

1. Complete the following table for the neutral isotopes below:

Isotope Name	Isotope Symbol	Atomic #	Mass #	# protons	# electrons	# neutrons
		15	31			
					27	31
Cobalt – 60						
	$^{208}_{85}\text{At}$					

2. Rubidium has two naturally-occurring isotopes: Rb-85 (isotopic mass = 84.912 u) and Rb-87 (isotopic mass = 86.909 u). Calculate the average atomic mass of Rubidium if Rubidium-85 makes up 72.17% of natural Rubidium.
(Compare your result to the value given on the periodic table.)
3. Naturally occurring bromine is composed of two isotopes, Br-79 with an isotopic mass of 78.918 amu, and Br-81, with an isotopic mass of 80.916 amu. Calculate the percent abundance of both Bromine isotopes. Report the final answers with 4 significant figures.